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From partial molar results (eq. (8.15)), the free energy of the solution is $G_{\text{soln}} = n_A G_{A0} + n_B G_{B0} + RT \ln n_A \text{Log}@XA D + n_B \text{Log}@XB D + HnA + nB L \text{Log}@PDL GA0 nA + GB0 nB + RT H\text{Log}@XA D nA + \text{Log}@XB D nB + \text{Log}@PD HnA + nBLL$ Because $GA0$ and $GB0$ are for constants that do not depend on subsequent increase in temperature, we need to solve for an increase in G_{soln} using Solve @ H1 é 2L RT H nA Log@XA D + nB Log@XB DL == RT H nA Log@XA D + nB Log@XB D + HnA + nB L Log@PDL , PD 99P Ø E -Log@XA D nA ...

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Single-Component Phase Diagrams (cont.): Thermodynamics of Solutions: Gaskell. "The free energy of solution." Chapter 11.5 in Introduction to Metallurgical Thermodynamics. pp. 328-338 E&R: 9.6 and 9.7. L17: X-ray Diffraction: A&T: Study 3.2.2 up to p. 140, and 3.4. Free Energy of Multi-phase Solutions at Equilibrium: E&R: 9.6 and 9.7. L18

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